

Group 14 Elements:-The Carbon Family

Group 14 includes carbon (C), silicon (Si), germanium (Ge), tin (Sn) and lead (Pb).

General electronic configuration of carbon family is ns^2np^2 .

Covalent radius:-Covalent radius expected to increase from C to Si,

From Si to Pb small increase is found.

Ionization Enthalpy:-The first ionization enthalpies of group 14 elements are higher than those of the corresponding group 13 elements.

Electronegativity:-Group 14 elements are smaller in size as compared to group 13 elements that's why this group elements are slightly more electronegative than group 13

Chemical properties:-

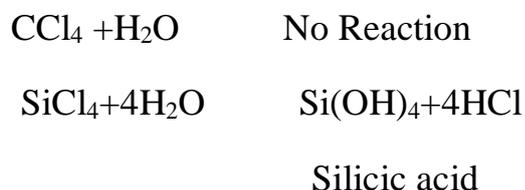
Carbon and silicon mostly show +4 oxidation state. Germanium forms stable compounds in +4 state and only few compounds in +2 state.

Tin forms compounds in both oxidation states. Lead compounds in +2 state are stable and in +4 state are strong oxidizing agents.

Exception:- Pb_4 and SnF_4 are ionic in nature.

Except CCl_4 other tetrachlorides are easily hydrolysed by water.

Since carbon does not have d-orbitals and hence cannot expand its coordination number beyond 4



Allotropes of Carbon:-The three types of allotropes are –

1-Diamond

2-Graphite

3-Fullerene

Diamond:-In diamond each carbon atom undergoes SP^3 hybridisation.

Each carbon is tetrahedrally linked to four other carbon atoms.

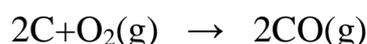
Graphite:-In graphite, carbon is SP^2 -hybridized. Graphite has a two-dimensional sheet like structure consisting of a number of hexagonal rings fused together.

Graphite conducts electricity along the sheet. It is very soft and slippery.

Fullerene Fullerene was discovered collectively by three scientists namely R.E Smalley, R.F Curl and H.W Kroto

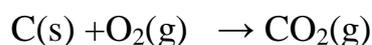
SOME Important Compounds Of Carbon and Silicon

Carbon monoxide:-It is prepared by direct oxidation of C in limited supply of oxygen.



Commercially it is prepared by the passage of steam over hot coke

Carbon Dioxide:-It is prepared by complete combustion of carbon and carbon fuels in excess of air.



Laboratory method:-

In laboratory it is prepared by the treatment of dil HCl on $CaCO_3$



Silicon dioxide:-Silicon dioxide is a COVALENT THREE DIMENSIONAL NETWORK SOLID.

Each silicon atom is covalently bonded in a tetrahedral manner to four oxygen atoms.

Silicones:-Silicones are the synthetic organo-silicon polymers having general formulae $(R_2SiO)_n$ in which R = alkyl (methyl, ethyl or phenyl)

Silicates:-Silicates exist in nature in the form of feldspar, zeolites, mica and asbestos etc.

The basic structure of silicates is SiO_4^{4-}

Zeolites:-Zeolites is a alumino-silicate of metal. Metal cations participating in formation of Zeolite are use usually Na^+ , K^+ , or Ca^{2+} .

Zeolites are used to remove permanent hardness of water.

ONE MARK QUESTIONS

1. Why is boron used in nuclear reactions?

Ans:-Because Boron can absorb neutrons.

2. By giving a balanced equation show how $\text{B}(\text{OH})_3$ behaves as an acid in water.



3. Name the element of group 14 which exhibits maximum tendency for catenation?

Ans:-Carbon

4. What is the basic building unit of all silicates?

Ans:- SiO_4^{4-} is the basic unit of all silicates.

5. What happens when NaBH_4 reacts with iodine?



6. What happens when boric acid is heated



7. What is producer gas?

Ans:-Producer gas is a mixture of CO and N_2 in the ratio of 2:1.

8. Write the state of hybridization of 'B' in BF_3 .

ANS:-Hybridisation of 'B' in BF_3 is Sp^2 .

9. Mention the state of hybridization in B in BH_4^- .

Ans:- Sp^3 .

10. Which oxide of carbon is regarded as anhydride of carbonic acid.

Ans:- CO_2 is regarded as a hydride of carbonic acid .



TWO MARKS QUESTIONS

1. Give the chemical reaction as an evidence for each of the following observations.

(i) Tin (II) is a reducing agent whereas lead (II) is not.

(ii) Gallium (I) undergoes disproportionation reaction.

Ans:- (i) Due to inert pair effect Pb^{2+} is more stable than Pb^{4+} . Whereas Sn^{4+} is more stable than Sn^{2+} .



This is because Ga^{3+} is more stable than Ga^+ .

2. What happens when

(i) Quick lime is heated with coke?

(ii) Carbon monoxide reacts with Cl_2



3. Give reason

(i) C and Si are always tetravalent but Ge, Sn, Pb show divalency.

(ii) Gallium has higher ionization enthalpy than Al. Explain.

Ans:- (i) Ge, Sn, Pb show divalency due to inert pair effect, Pb^{2+} is more stable than Pb^{4+} .

(ii) Due to poor shielding effect of d-electrons in Ga effective nuclear charge increases as compared to Al thus the I.E is higher than Al.

4. Give reason why boron and aluminium tend to form covalent compounds.

Ans:- Sum of three ionization of both the elements are very high. Thus they have no tendency to lose electrons to form ionic compounds. Instead they form covalent compounds.

5. If B-Cl bond has a dipole moment, Explain why BCl_3 molecule has zero dipole moment.

Ans:- B-Cl bond has dipole moment because of polarity. In BCl_3 since the molecule is symmetrical thus the polarities cancel out.

6. Suggest a reason as to why CO is poisonous.

Ans:- CO reacts with haemoglobin to form carboxy-haemoglobin which can destroy the oxygen carrying capacity of haemoglobin and the man dies of suffocation.

7. What do you understand by-

(a) Inert pair effect:- The pair of electron in the valence shell does not take part in bond formation it is called inert pair effect.

(b) Allotropy:- It is the property of the element by which an element can exist in two forms which have same chemical properties but different physical properties due to their structures.

8. How is excessive content of CO_2 responsible for global warming?

Ans:- Excess of CO_2 absorbs heat radiated by the earth. Some of it is dissipated into the atmosphere while the remaining part is radiated back to the earth. Temperature of the earth increases.

9. Describe two similarities and two dissimilarities between B and Al.

Ans:- Similarities:-

- (i) Both have same number of valence electrons.
- (ii) Both have similar electronic configuration.

Dissimilarities:-

- (i) B is a non-metal where Al is a metal
- (ii) B forms acidic oxide whereas Al forms amphoteric oxides.

10. What are fullerenes? How were they prepared?

Ans:- Fullerenes are the allotropes of carbon. Its structure is like a soccer ball.

They are prepared by heating graphite in electric arc in presence of inert gases such as helium or argon.

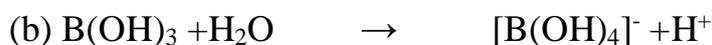
THREE MARKS QUESTIONS

1. What happens when

- (a) Borax is heated strongly

(b) Boric acid is added to water

(c) Aluminium is treated with dilute NaOH



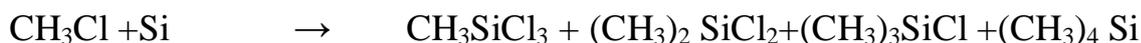
2. Explain the following reactions.

(a) Silicon is heated with methyl chloride at high temperature in the presence of copper.

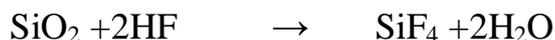
(b) Silicon dioxide is treated with hydrogen fluoride.

(c) CO is heated with ZnO.

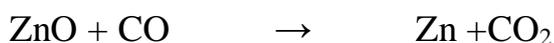
Ans:- (a) A mixture of mono-, di- and trimethylchlorosilanes along with a small amount of tetramethylsilane is formed.



(b) The initially formed silicon tetrafluoride dissolves in HF to form hydrofluorosilicic acid



(c) ZnO is reduced to zinc metal



3. Give reasons:-

(a) Diamond is used as an abrasive.

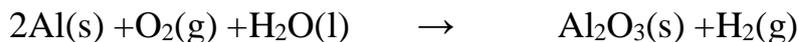
(b) Aluminium alloys are used to make aircraft body.

(c) Aluminium utensils should not be kept in water overnight.

Ans:- (a) Diamond is used as an abrasive because it is an extremely hard substance.

(b) Alloys of aluminium like duralium is used to make aircraft body due to some of its properties.

(c) Generally aluminium metal does not react with water quickly but when it is kept overnight. It reacts slowly with water in presence of air.

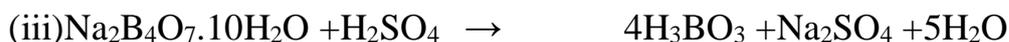


4. A certain salt X, gives the following results.

(i) Its aqueous solution is alkaline to litmus.

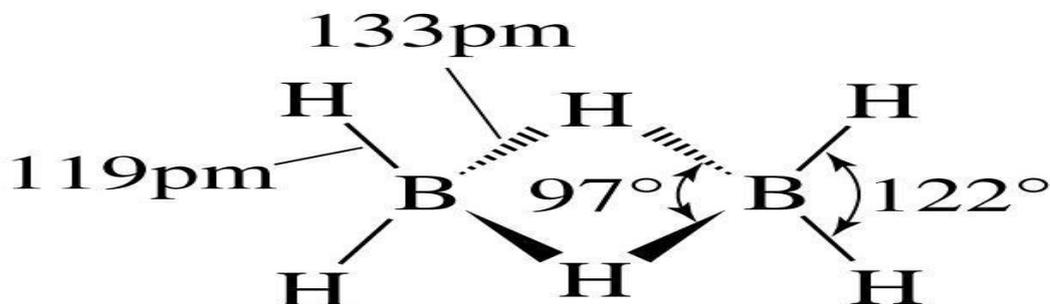
(ii) It swells up to a glassy material Y on strong heating.

(iii) when conc. H_2SO_4 is added to a hot solution of X, white crystal of an acid Z separates out.



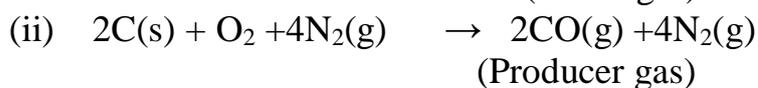
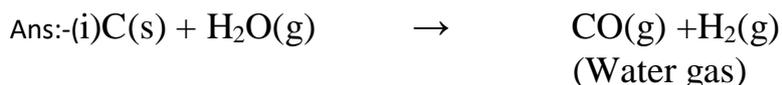
5. draw structure of diborane .

Ans.



FIVE MARKS QUESTIONS

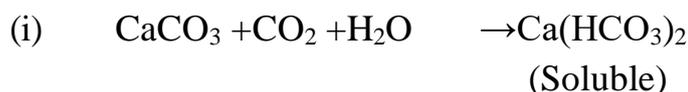
1 Explain the formation of (i) Water gas (ii) Producer gas. Give their uses. What happens when CO_2 is passed through lime water (i) for short duration (ii) for long duration.



Water gas and Producer gas are used as fuel.



(White ppt.)



- 2 (a) Why do Boron halides form addition compound with NH_3 ?
(b) Assign appropriate reason for each of the following observations :-

- (i) Anhydrous AlCl_3 is used as a catalyst in many organic reactions.
(ii) No form of elemental silicon is comparable to graphite.

Ans:- (a) It is because BX_3 is electron deficient whereas NH_3 is electron rich.

(b) (i) It is Lewis acid.

(ii) It cannot form $p\pi - p\pi$ bond due to large size.

3. (i) Give reason for the following observations:-
(a) The tendency for catenation decreases down the group in Group 14.
(b) The decreasing stability of +3 oxidation state with increasing atomic number in group 13.
(c) PbO_2 is a stronger oxidizing agent than SnO_2 .
(d) Molten aluminium bromide is a poor conductor of electricity.

Ans:- (i)(a) It is due to decrease in bond dissociation energy which is due to increase in atomic size.

$\text{C-C} > \text{Si-Si} > \text{Ge-Ge} > \text{Sn-Sn} > \text{Pb-Pb}$.

(b) It is due to inert pair effect.

(c) PbO_2 is stronger oxidizing agent than SnO_2 because Pb^{2+} is more stable than Pb^{4+} whereas Sn^{4+} is more stable than Sn^{2+} .

(d) Molten AlBr_3 is poor conductor of electricity because it is covalent compound.