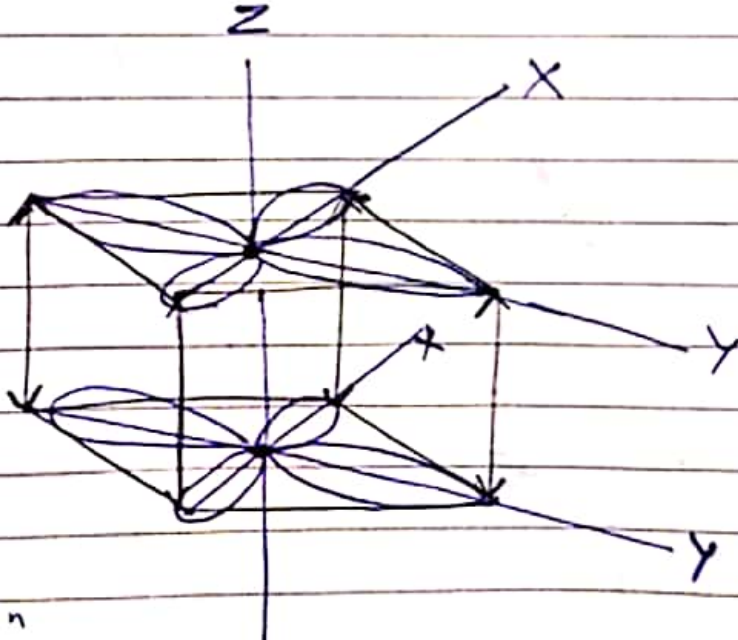
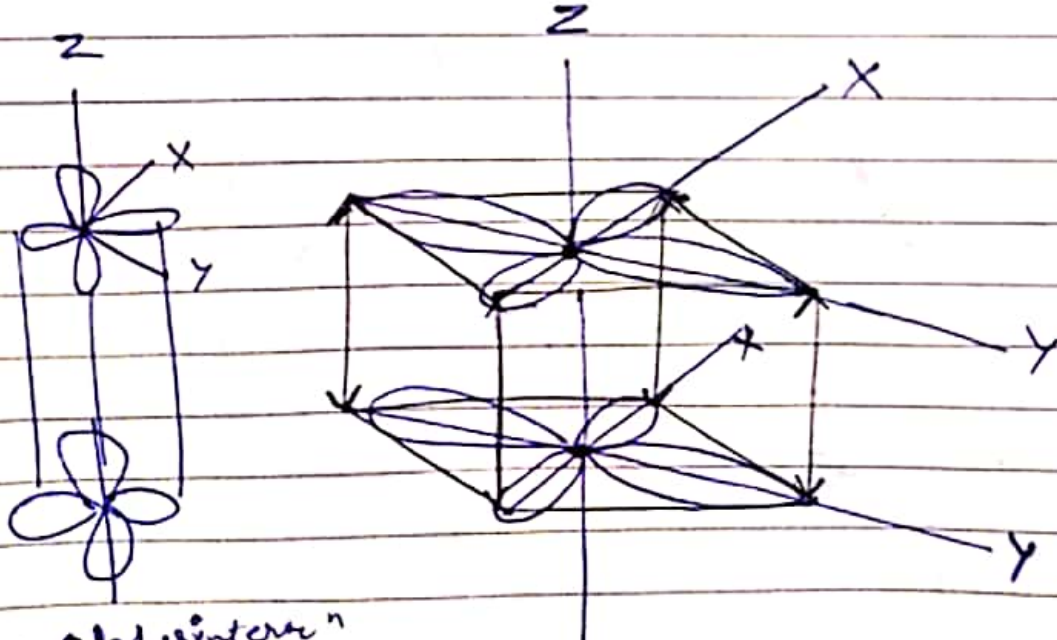
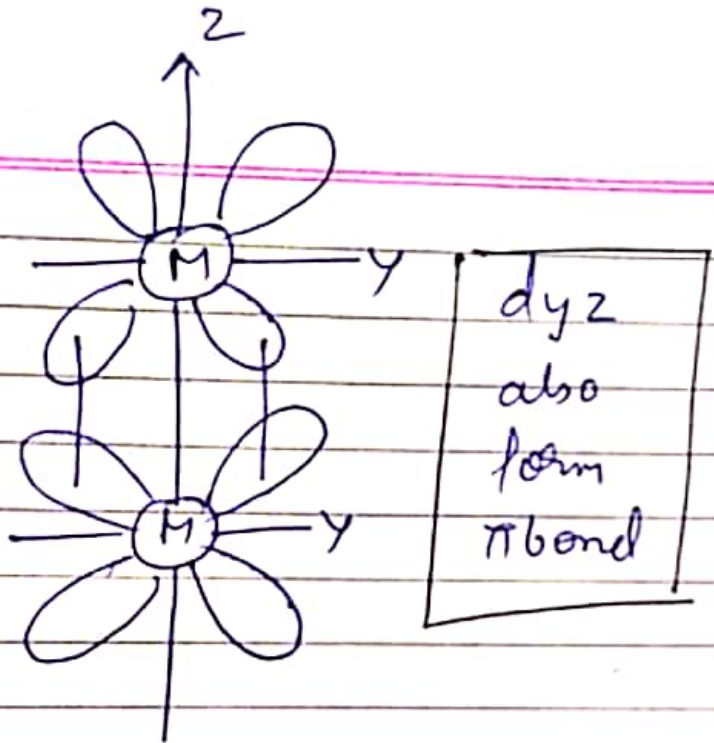


Delta bond is formed by d orbitals because in delta bond four lobe interaction is there . it is stronger than pie bond because 4 lobes are interacting here.

Maximum delta bond between 2 metal can be 2 .
Therefore maximum bonds which can be form between 2 metal is 5(1 sigma ,2pie , 2 delta bonds)

If we consider z axis as internuclear axis between 2 metals than (dx^2-y^2 , dxy) will form delta bond because 4 lobes interaction will be there in that case and dz^2 will form sigma bond and dxz and dyz will form pie bond. Example in $(Re_2Cl_8)^{2-}$ one delta bond is found.



$dx^2 - y^2$ form
s bond
bleakht let
as sum inter-
nuclear axis as
z axis.

∴ this is
interaction b/w floras