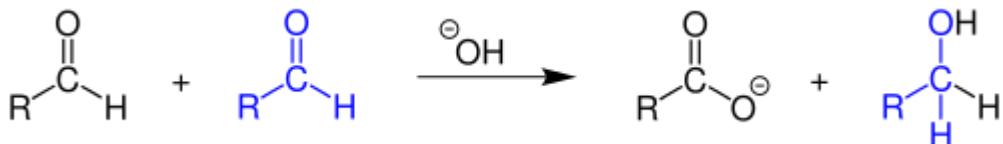


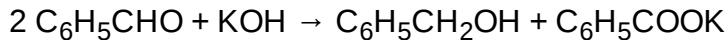
Cannizzaro reaction

The **Cannizzaro reaction**, named after its discoverer Stanislao Cannizzaro, is a chemical reaction which involves the base-induced disproportionation of two molecules of a non-enolizable aldehyde to give a primary alcohol and a carboxylic acid.^{[1][2]}

Cannizzaro reaction	
Named after	Stanislao Cannizzaro
Reaction type	Organic redox reaction
Identifiers	
Organic Chemistry Portal	cannizzaro-reaction
RSC ontology ID	RXNO:0000218



Cannizzaro first accomplished this transformation in 1853, when he obtained benzyl alcohol and potassium benzoate from the treatment of benzaldehyde with potash (potassium carbonate). More typically, the reaction would be conducted with sodium hydroxide or potassium hydroxide, giving the sodium or potassium carboxylate salt of the carboxylic-acid product:



The process is a redox reaction involving transfer of a hydride from one substrate molecule to the other: one aldehyde is oxidized to form the acid, the other is reduced to form the alcohol.^[3]

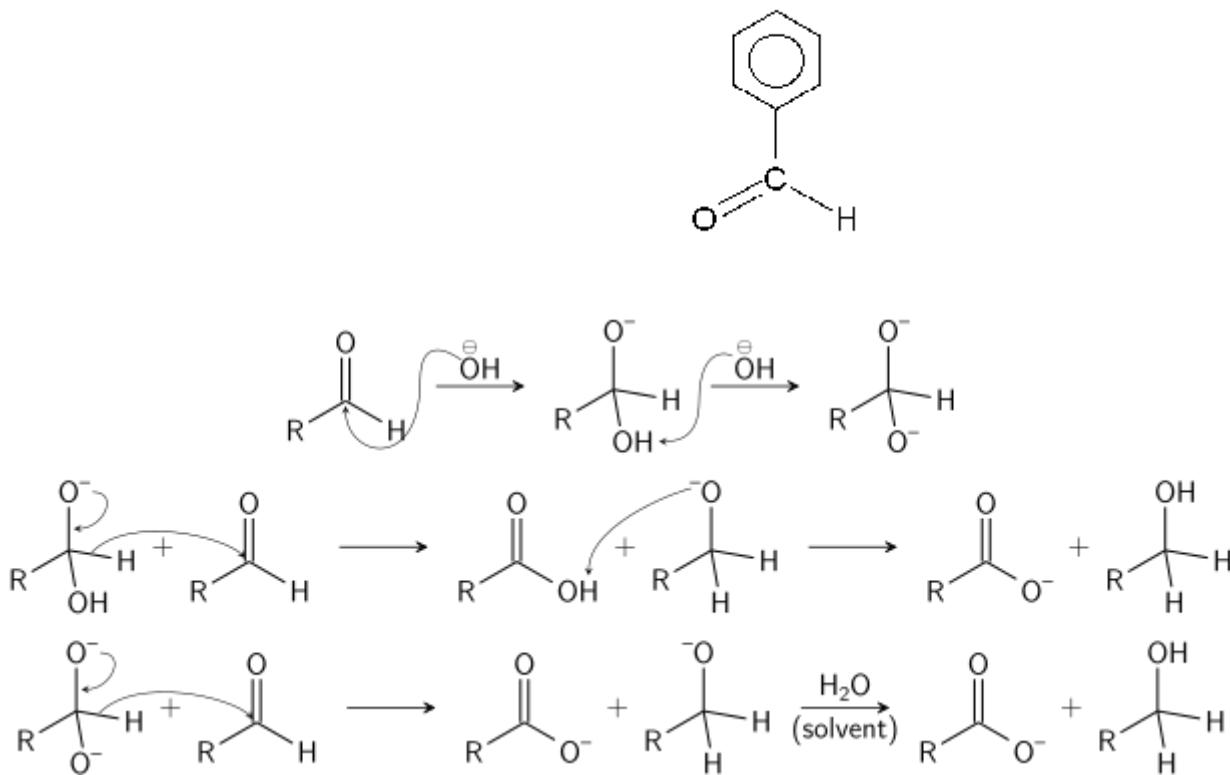
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Mechanism

The reaction involves a nucleophilic acyl substitution on an aldehyde, with the leaving group concurrently attacking another aldehyde in the second step. First, hydroxide attacks a carbonyl. The resulting tetrahedral intermediate then collapses, re-forming the carbonyl and transferring hydride to attack another carbonyl.^[4]

In the final step of the reaction, the acid and alkoxide ions formed exchange a proton. In the presence of a very high concentration of base, the aldehyde first forms a doubly charged anion from which a hydride ion is transferred to the second molecule of aldehyde to form carboxylate and alkoxide ions. Subsequently, the alkoxide ion acquires a proton from the solvent.



Overall, the reaction follows third-order kinetics. It is second order in aldehyde and first order in base:

$$\text{rate} = k[\text{RCHO}]^2[\text{OH}^-]$$

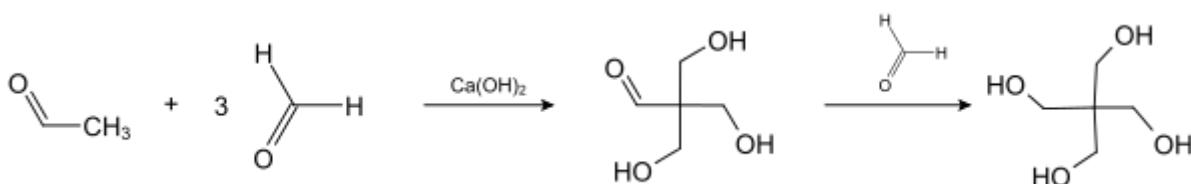
At very high base a second path (k') becomes important that is second order in base:

$$\text{rate} = k[\text{RCHO}]^2[\text{OH}^-] + k'[\text{RCHO}]^2[\text{OH}^-]^2$$

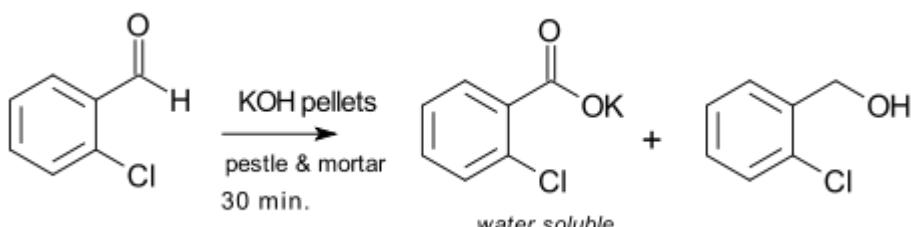
The k' pathway implicates a reaction between the doubly charged anion (RCHO_2^{2-}) and the aldehyde. The direct transfer of hydride ion is evident from the observation that the recovered alcohol does not contain any deuterium attached to the α -carbon when the reaction is performed in the presence of D_2O .

Scope

Due to the strongly alkaline reaction conditions, aldehydes that have alpha hydrogen atom(s) instead undergo deprotonation there, leading to enolates and possible aldol reactions. Under ideal conditions the reaction produces 50% of both the alcohol and the carboxylic acid (it takes two aldehydes to produce one acid and one alcohol).^[5] This can be economically viable if the products can be separated and both have a value; the commercial conversion of furfural into furfuryl alcohol and 2-furoic acid is an example of this.^[6] Alternatively, higher yields of one product (usually the alcohol) can be achieved in the **crossed Cannizzaro reaction**, in which a sacrificial aldehyde is used in combination with a more valuable chemical. In this variation, the reductant is formaldehyde, which is oxidized to sodium formate and the other aldehyde chemical is reduced to the alcohol. Thus, the yield of the valuable chemical is high, although the atom economy can be low. The final stage in the synthesis of pentaerythritol is an example.

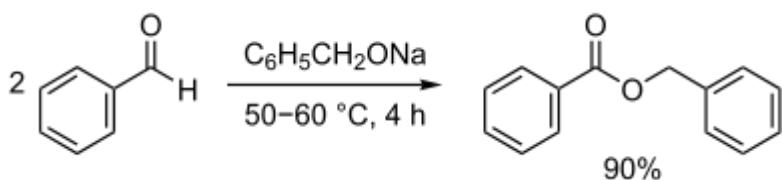


A solvent-free reaction has been reported involving grinding liquid 2-chlorobenzaldehyde with potassium hydroxide in a mortar and pestle:^[7]



Variations

In the Tishchenko reaction, the base used is an alkoxide rather than hydroxide, and the product is an ester rather than the separate alcohol and carboxylate groups. After the nucleophilic base attacks an aldehyde, the resulting new oxygen anion attacks another aldehyde to give a hemiacetal linkage between two of the formerly aldehyde-containing reactants rather than undergoing tetrahedral collapse. Eventually tetrahedral collapse does occur, giving the stable ester product.



Certain ketones can undergo a Cannizzaro-type reaction, transferring one of their two carbon groups rather than the hydride that would be present on an aldehyde.^[8]

See also

- Meerwein–Ponndorf–Verley reduction
- Oppenauer oxidation

References

1. Cannizzaro, S. (1853). "Ueber den der Benzoësäure entsprechenden Alkohol" (<https://babel.hathitrust.org/cgi/pt?id=chi.44985506;view=1up;seq=523>) [On the alcohol corresponding to benzoic acid]. *Liebigs Annalen der Chemie und Pharmacie*. **88**: 129–130. doi:[10.1002/jlac.18530880114](https://doi.org/10.1002/jlac.18530880114) (<https://doi.org/10.1002%2Fjac.18530880114>).
2. List, K.; Limpricht, H. (1854). "Ueber das sogenannte Benzoëoxyd und einige andere gepaarte Verbindungen" (<https://babel.hathitrust.org/cgi/pt?id=uva.x002457941;view=1up;seq=586>) [On so-called benzoic oxide and some other paired compounds]. *Liebigs Annalen der Chemie und Pharmacie*. **90** (2): 190–210. doi:[10.1002/jlac.18540900211](https://doi.org/10.1002/jlac.18540900211) (<https://doi.org/10.1002%2Fjac.18540900211>).

3. Geissman, T. A. "The Cannizzaro Reaction" *Org. React.* **1944**, 2, 94.
[doi:10.1002/0471264180.or002.03](https://doi.org/10.1002/0471264180.or002.03) (<https://doi.org/10.1002%2F0471264180.or002.03>)
(Review)
4. Smith, Michael B.; March, Jerry (2007), *Advanced Organic Chemistry: Reactions, Mechanisms, and Structure* (<https://books.google.com/books?id=JDR-nZpojECC&printsec=frontcover>) (6th ed.), New York: Wiley-Interscience, ISBN 978-0-471-72091-1
5. W. C. Wilson (1941). "2-Furancarboxylic Acid and 2-Furylcarbinol" (<http://www.orgsyn.org/demo.aspx?prep=CV1P0276>). *Organic Syntheses*; *Collective Volume*, **1**, p. 276
6. Mariscal, R.; Maireles-Torres, P.; Ojeda, M.; Sádaba, I.; López Granados, M. (2016). "Furfural: a renewable and versatile platform molecule for the synthesis of chemicals and fuels" (https://digital.csic.es/bitstream/10261/184700/1/Mariscal_Furfural-A%20renewable_versatile_2016_postprint.pdf) (PDF). *Energy Environ. Sci.* **9** (4): 1144–1189.
doi:10.1039/C5EE02666K (<https://doi.org/10.1039%2FC5EE02666K>). hdl:10261/184700 (<https://hdl.handle.net/10261%2F184700>). ISSN 1754-5692 (<https://www.worldcat.org/issn/1754-5692>).
7. A Facile Solvent-Free Cannizzaro Reaction Phonchaiya, Sonthi; Panijpan, Bhinyo Rajviroongit, Shuleewan; Wright, Tony; Blanchfield, Joanne T. "A Facile Solvent-Free Cannizzaro Reaction" *J. Chem. Educ.* **2009**, volume 86, page 85. doi:10.1021/ed086p85 (<https://doi.org/10.1021%2Fed086p85>)
8. Jiang, Xin-Dong; Matsukawa, Shiro; Kakuda, Ken-ichiro; Fukuzaki, Yuta; Zhao, Wei-Li; Li, Lin-Song; Shen, Huai-Bin; Kojim, Satoshi; Yamamoto, Yohsuke (2010). "Efficient synthesis of tetradecafluoro-4-phenylheptan-4-ol by a Cannizzaro-type reaction and application of the alcohol as a bulky Martin ligand variant for a new anti-apicophilic phosphorane". *Dalton Trans.* **39** (41): 9823–9829. doi:10.1039/C0DT00539H (<https://doi.org/10.1039%2FC0DT00539H>). PMID 20859600 (<https://pubmed.ncbi.nlm.nih.gov/20859600/>).

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