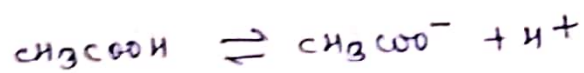


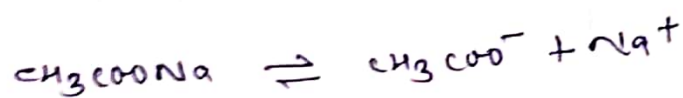
Common Ion Effect :-

If a salt of a weak acid is added to a soln of the acid itself, the dissociation of the acid is diminished further.

for example-

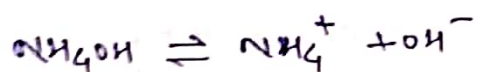


The addition of sodium acetate to a soln of acetic acid suppresses the dissociation of acetic acid.

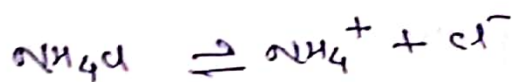


The addition of one of the products of dissociation (acetate ion) supplied by the largely dissociated salt (sod. acetate) pushes the equilibrium to the left. Thus the dissociation of acetic acid is suppressed.

Likewise, the dissociation of a weak base, such as NH_4OH



is suppressed on the addition of a salt like NH_4Cl which supplies NH_4^+ ion.



The addition of strong base such as ~~NaOH~~ NaOH also suppresses the dissociation of NH_4OH .

Thus, the suppression of the dissociation of a weak acid or weak base on the addition of its own ions is called common ion effect.